

## Analyzing the Bohr Atom

Use the information supplied in the table to fill in the remaining blanks in each row.

Number of Electrons	Electron Configuration	Mass Number	Number of Protons	Number of Neutrons	Element	Symbol	Isotope Notation
5	2-3	11	5	6	Boron	B	<sup>11</sup> <sub>5</sub> B
9	2-7	19	9	10	Fluorine	F	<sup>19</sup> <sub>9</sub> F
13	2-8-3	27	13	14	aluminum	Al	<sup>27</sup> <sub>13</sub> Al
6	2-4	12	6	6	carbon	C	<sup>12</sup> <sub>6</sub> C
11	2-8-1	23	11	12	Sodium	Na	<sup>23</sup> <sub>11</sub> Na
1	1	1	1	0	Hydrogen	H	<sup>1</sup> <sub>1</sub> H
19	2-8-8-1	39	19	20	Potassium	K	<sup>39</sup> <sub>19</sub> K
17	2-8-7	35	17	18	chlorine	Cl	<sup>35</sup> <sub>17</sub> Cl
8	2-6	16	8	8	oxygen	O	<sup>16</sup> <sub>8</sub> O
20	2-8-8-2	40	20	20	Calcium	Ca	<sup>40</sup> <sub>20</sub> Ca
16	2-8-6	32	16	16	sulfur	S	<sup>32</sup> <sub>16</sub> S
18	2-8-8	40	18	22	Argon	Ar	<sup>40</sup> <sub>18</sub> Ar
4	2-2	9	4	5	Beryllium	Be	<sup>9</sup> <sub>4</sub> Be
2	2	4	2	2	Helium	He	<sup>4</sup> <sub>2</sub> He
21	2-8-9-2	45	21	24	Scandium	Sc	<sup>45</sup> <sub>21</sub> Sc